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particularly efficiently if the two components, under certain first conditions, for example under certain pressure and temperature conditions, form a solubility gap in the concentrations that are present.

5 For the purposes of definition, in the following "mixture" means a system comprising two or more types of molecules, the chemical and physical characteristics of which are spatially constant (homogeneous system). A solution is a mixture with which one material or one type of molecule is present in excess. Two liquids form a solubility gap if they  
10 cannot be mixed together without limitation; one then obtains two liquid phases in which the components of the liquid are present in a varying composition, for example, the one component extensively in the one phase and the other component extensively in the other phase. A solubility gap can be observed in that the clear liquid becomes turbid  
15 with a change in temperature, i.e. forms an emulsion, that is an indication of the phase separation. However, the turbidity or emulsion is not a necessary indication for a solubility gap; there are so-called micro emulsions in which the two phases are finely distributed such that the liquid continues to be optically clear.

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Due to phenomena that up to now are not fully understood, in general a liquid that is composed of two components and is in the state of a

solubility gap has a better cleaning capacity than do the two individual components if they are used in a pure state or in a highly concentrated state one after the other. Perhaps the excellent cleaning effect of liquids that are present in a solubility gap is brought about by interaction at the interfaces between the two phases, and possibly additionally by mechanical effects due to the droplets that are frequently held in distinct movement via ultrasound or a stirring mechanism. The use of the liquid in the state of the solubility gap thus enables an advantageous cleaning not only with regard to its cleaning effect but also with regard to the duration and with regard to the quantities of individual components that are required.

So that the cleaning liquid can remain in use as long as possible, it must be freed from the contaminations that it has received. Pursuant to the invention, this takes place in that the cleaning liquid is brought out of the state of the solubility gap into the state of a true mixture, i.e. a homogeneous state. From this homogeneous liquid the contaminations can be removed, depending on the nature thereof, via a filter (especially inorganic pigment-containing contaminations), or in that the contaminations, as a consequence of their densities that are different from the liquid, accumulate at the base or on the surface of the liquid, from where they are withdrawn (especially fatty

contaminations). If a filtration or a separation of the liquid is, in contrast, to be effected in the state of the solubility gap, in so doing also a large percentage of at least that component is separated off that ties up or captures a respective contamination or binds it to its contact surface.

On the whole, as a result of the planned conversion of the cleaning liquid on the one hand into the state of the solubility gap for the cleaning, and on the other hand into the state of the true mixture for the separation of the contaminations, an efficient method is provided for the liquid cleaning of objects, which method, with an extensive ability to recycle the cleaning liquid (separation of contaminations), enables an effective cleaning of very different types of objects. It is to be understood that the composition of the cleaning liquid is adapted to the respective cleaning problem, whereby it is merely mandatory to select such components that under first designated conditions form a solubility gap, and under second designated conditions mix with one another.

With the features of claim 2 there is achieved the advantage that not only inorganic but also organic dirt can be dissolved, whereby the organic component can frequently be present in a relatively low

concentration and yet cleans as if it were present in a higher concentration.

5 A very good cleaning effect for a wide variety of types of contamination is achieved with the features of claim 3, whereby numerous organic components that are provided with molecules having lipophilic and hydrophilic groups form a solubility gap with water.

10 With the features of claim 4, a cleaning method is provided according to which the cleaning liquid comprises predominantly water.

15 The method of the present invention can be carried out in a particularly straightforward manner if the state of the solubility gap changes over into the state of the mixture by merely altering the temperature. Other possibilities for converting the two states into one another comprise a change in pressure, a particularly intensive agitation, e.g. by means of ultrasound, by introduced contaminations that lead to a shifting of an equilibrium or to an unstable state suddenly changing over into a stable one, etc.

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The method pursuant to claim 6 is particularly advantageous, since the cleaning effect is generally better at a higher temperature than at a lower temperature.

5 Claim 7 characterizes an embodiment of the method that is particularly effective relative to the separation of dirt from the cleaning liquid.

The inventive method is particularly suitable for all liquid cleanings where no chemical reaction takes place between the contamination and the cleaning liquid that varies the molecular composition of the cleaning liquid. The cleaning liquid in the state of the solubility gap is a medium with which contaminations are effectively transferred from the uncleaned surface into the cleaning liquid. The conversion of the cleaning liquid from the state of the solubility gap into the state of the homogeneous mixture is the key for being able to effectively remove the contaminations contained in the cleaning liquid from the liquid.

The invention will subsequently be explained with the aid of one example and the accompanying single figure, which illustrates an apparatus for carrying out the inventive method.

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straight chained and branched alkyl groups having up to 18  
C atoms

5 In the described example, the cleaning liquid contains 90 wt.-% water  
and 10 wt.-% glycol ether, preferably dipropylene glycol mono-n-  
propyl ether.

10 The designated cleaning liquid is contained in a cleaning tank 2, from  
which a line 6, which is provided with a pump 4 for controlling the flow  
velocity, leads into a separation tank 8. The separation tank 8 is  
connected via an overflow 9 with a collection tank 10, from which a  
return line 14, which is provided with a pump 12, leads through a filter  
device 16 back to the cleaning tank 2. Contained in the cleaning tank  
2 is an agitation device 16, for example a stirring mechanism and/or an  
15 ultrasound device. Each of the tanks 2, 8 and 10 is provided with its  
own tempering or temperature control device 18 by means of which the  
temperature of the tanks can be held at a predetermined value  
independently of one another. Disposed over the cleaning tank 2 is a  
transport mechanism 20 for receiving the objects that are to be  
20 cleaned.

The function of the described apparatus, which on the whole operates at atmospheric pressure, is as follows:

5 The previously described cleaning liquid is optically clear at room temperature, i.e. the organic component forms a true mixture with the water. If the cleaning liquid is heated to 40°, a turbidity occurs, which indicates that the solubility of the organic component in the water is exceeded and a two-phase system forms, with organic-rich droplets in a continuous aqueous phase. The cleaning tank 2 is held at a  
10 temperature of 40°, and the cleaning liquid disposed therein is intensely swirled with the agitation device 16. The transport mechanism 20 is lowered into the cleaning tank 2, so that the objects that are to be cleaned come into intensive contact with the cleaning liquid, which is in the state of the solubility gap. In so doing, the  
15 cleaning liquid is continuously pumped off into the separation tank 8 via the pump 4, with the separation tank being held at a temperature of only 20°, so that the contaminated cleaning liquid is present at that location in the state of the true mixture. Organic dirt, which is specifically lighter than the liquid, is deposited on the surface and can  
20 be removed with a rake 22 or some other device. Specifically heavier dirt is deposited at the base of the separation tank 8, where it can be withdrawn via a non-illustrated known device.

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As is clear from the preceding, the cleaning liquid serves as a transport medium for the contaminations by removing and receiving these contaminations from the objects in the cleaning tank 2, subsequently giving up the contaminations in the separation tank 8 by separation and giving up the contaminations in the filter device by filtration.

The described system can be modified such that for example in a utensil rinsing machine or washing machine in the cleaning tank, the described method occurs, with the cleaning liquid then being pumped out of the cleaning tank into a storage tank where it is stored, while in the cleaning tank only rinsing procedures take place. The cleaning liquid can subsequently be pumped back into the cleaning tank for the cleaning of further objects. In this manner, the cleaning liquid can be used many times for the cleaning of objects, and need only occasionally be topped off. The separated contaminations can be removed with the rinsed water.

Further examples for the basic composition of liquids having a solubility gap, which composition comprises water and an organic component, will be provided in the following. In this connection, in each case first the chemical designation of the organic component is provided, then the concentration up to which the organic component is soluble with

water at room temperature, and then the concentration up to which water can be added and be soluble with the organic component. In the first example, glycol ether, the solubility gap at room temperature is thus between 5% and 82% glycol ether in 95% or 18% water respectively. The subsequent information (first example): MPC (Multi Phase Cleaning) at 5% starting at 29° C indicates respectively the concentration of the organic component with which one advantageously operates with the respective liquid, and the temperature above which a good cleaning effect is achieved due to the stable solubility gap. For the complete mixing or for the separation of the contaminations, the liquid is respectively advantageously cooled off to room temperature. It is to be understood that one advantageously works with concentrations that are slightly, e.g. 0.1 to 0.2 %, below the concentration at which the solubility gap occurs at room temperature.

Glycol Ether:

- Propyleneglycolmonobutylether PnB

Water Solubility      5%

Water in PnB      18%

MPC at 5% starting at 29° C

- Dipropyleneglycolmono n-butylether DPnB

Water Solubility      4%

Water in DPnB 14%

MPC at 1% starting at 23°C

- Tripropyleneglycolmono n-butylether TPnB

5 Water Solubility 3%

Water in TPnB 8%

MPC at 3% starting at 23° C

- Tripropyleneglycolmonopropylether TPnP

10 Water Solubility 5%

Water in TPnP 12%

MPC at 5% starting at 45° C

- Propyleneglycolphenylether PPH

15 Water Solubility 1%

Water in PPH 6%

MPC at 1% starting at 23° C

#### Ester / Acetate

20 - Propyleneglycoldiacetate PGDA

Water Solubility 8%

Water in PGDA 4%

MPC at 8% starting at 23° C

- Dipropyleneglycolmonoethyletheracetate DPMA

Water Solubility 5%

Water in DPMA 8%

MPC at 5% starting at 42° C

- Butylglycolacetate

Water Solubility 1.5%

Water in Butylglycolacetate 5%

MPC at 1.5% starting at 30° C

- Butyldiglycolacetate

Water Solubility 6.5%

Water in Butyldiglycolacetate 10%

MPC at 6.5% starting at 35° C

### Alcohols

- Cyclohexanol

Water Solubility 3.7%

Water in Cyclohexanol 7%

MPC at 3.7% starting at 35° C

- Hexylalcohol

Water Solubility 0.6%

5 Water in Hexylalcohol 0%

MPC at 10% starting at 35° C

-2-Ethyl-1-hexanol

Water Solubility 0.1%

10 Water in 2-Ethyl-1-hexanol 0%

MPC at 0.1% starting at 22° C